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INDIAN SCHOOL MUSCAT SECOND PERIODIC TEST

CHEMISTRY

CLASS: XII

Sub. Code: 042

Time Allotted: 50mts

16.09.2018

Max. Marks: 20

GENERAL INSTRUCTIONS:

- All questions are compulsory.
- Questions 1-3 are very short answer questions carries 1 mark each.
- Questions 4-7 are short answer questions carry 2 marks each.
- Questions 8-10 are short answer questions carry 3 marks each.

1. Draw the structure of XeO_3 molecule. 1
2. What happens when chlorine gas is passed through hot concentrated sodium hydroxide solution? 1
3. Of all noble gases, xenon forms well established compounds. Why? 1
4. Complete the following equations 2
 - a) $\text{XeF}_4 + \text{SbF}_5 \rightarrow$
 - b) $\text{SO}_2 + \text{MnO}_4^- + \text{H}_2\text{O} \rightarrow$
5. Give reason 2
 - a) Ammonia has higher boiling point than phosphine
 - b) ICl is more reactive than I_2
6. How is ozone estimated quantitatively? 2
7. Account for the following 2
 - a) Oxygen shows lower tendency for catenation than sulphur
 - b) Fluorine is stronger oxidizing agent than chlorine
8. Arrange the following in the increasing order of property mentioned in the bracket 3
 - a) NH_3 , PH_3 , SbH_3 , AsH_3 , BiH_3 [basic character]
 - b) H_2S , H_2O , H_2Se , H_2Te [acidic character]
 - c) HBr , HF , HCl , HI [reducing character]
9. Draw the structures of the following molecules 3
 - a) HClO_3
 - b) $\text{H}_2\text{S}_2\text{O}_8$
 - c) H_3PO_2
10. Explain with balanced equations the manufacture of sulphuric acid by contact process. Also mention the favorable conditions to maximize its yield. 3

End of the Question Paper